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Key indicators

Single-crystal synchrotron study
 $T = 123\text{ K}$
 Mean $\sigma(\text{C}-\text{C}) = 0.015\text{ \AA}$
 Disorder in solvent or counterion
 R factor = 0.089
 wR factor = 0.243
 Data-to-parameter ratio = 13.1

For details of how these key indicators were automatically derived from the article, see <http://journals.iucr.org/e>.

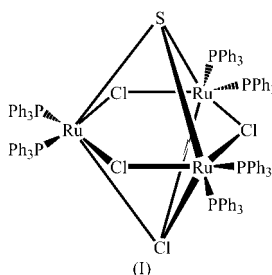
A synchrotron study of μ_3 -chloro-tri- μ -chloro- μ_3 -sulfido-hexakis(triphenylphosphine)triruthenium(II)–tetrahydrofuran–water (1/1.25/0.75)

The crystal structure of $[\text{Ru}_3(\mu_3\text{-Cl})(\mu\text{-Cl})_3(\mu_3\text{-S})(\text{PPh}_3)_6] \cdot 1.25\text{C}_4\text{H}_8\text{O} \cdot 0.75\text{H}_2\text{O}$, has been determined using synchrotron radiation. The cluster complex contains a central six-membered ring of alternating Ru and Cl atoms. The core is completed by capping chloride and sulfido groups.

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Comment

It is now well established that the reaction of metal halide complexes with $\text{S}[\text{Si}(\text{CH}_3)_3]_2$ produces metal sulfide clusters of high nuclearity. The ionic organometallic ruthenium sulfide cluster complexes $[\text{Ru}_3\text{S}_2(p\text{-cymene})_3]^{2+}$ (*p*-cymene is 4-isopropyltoluene) (Lockemeyer, Rauchfuss & Rheingold, 1989) and $[\{\text{RuCl}(\text{TMT})\}_3\text{S}]^+$ (TMT is tetramethylthiophene) (Lockemeyer, Rauchfuss, Rheingold & Wilson, 1989) have been prepared *via* this method from $[\text{RuCl}_2(p\text{-cymene})]_2$ and $[\text{RuCl}_2(\text{TMT})]_2$, respectively. The neutral ruthenium sulfide clusters $[\text{Ru}_4\text{S}_6(\text{PPh}_3)_4]$, $[\text{Ru}_5\text{S}_6(\text{PPh}_3)_5]$ and $[\text{Ru}_6\text{S}_8(\text{PPh}_3)_6]$ have been prepared from the reaction of $[\text{RuCl}_2(\text{PPh}_3)_3]$ with NaSH or $\text{S}[\text{Si}(\text{CH}_3)_3]_2$ under various conditions (Eckermann *et al.*, 2002).



Small crystals of the title compound, (I), were isolated from the reaction of $[\text{RuCl}_2(\text{PPh}_3)_3]$ with $\text{S}[\text{Si}(\text{CH}_3)_3]_2$; it is thought to be an intermediate in the formation of halide-free cluster compounds. The poorly diffracting crystals did not yield sufficient data on a conventional laboratory diffractometer. A single-crystal structure of (I) (Fig. 1) was eventually obtained using synchrotron radiation at the ChemMatCARS facility of the Advanced Photon Source.

The core of the cluster consists of a flat six-membered ring of alternating Ru atoms and bridging chloro ligands. Additional μ_3 -capping chloro and sulfido ligands, one on each side of the six-membered ring, complete the core. Each Ru centre is coordinated by two PPh_3 ligands and has distorted octahedral geometry with the bridging chloride occupying axial positions. Overall, the molecule has non-crystallographic C_{3v} symmetry. A similar core geometry has been observed in the cationic cluster $[\text{Ru}_3\text{Cl}_5((S)\text{-binap})_3]^+$ [binap is 2,2'-bis(diphenylphosphino)-1,1'-binaphthyl] (Mashima *et al.*, 1991),

which has two μ_3 -chlorides capping a six-membered Ru_3Cl_3 ring. The cation of the organometallic ruthenium–sulfide cluster $[\{(\text{TMT})\text{RuCl}_3\text{S}\}]^+$ also consists of a six-membered Ru_3Cl_3 ring which is capped by a μ_3 -S atom (Lockemeyer, Rauchfuss, Rheingold & Wilson, 1989). The bond distances of (I) are in accord with equivalent bond distances in these related clusters.

Experimental

The title compound was prepared from the reaction of $[\text{RuCl}_2(\text{PPh}_3)_3]$ and $\text{S}[\text{Si}(\text{CH}_3)_2]$ in refluxing tetrahydrofuran for 16 h. After this time, the reaction was cooled to room temperature, the solution filtered and the solvent volume reduced *in vacuo* (*ca* 5 ml). Microcrystals suitable for single-crystal X-ray diffraction analysis were grown from the reaction solution at 243 K.

Crystal data

$[\text{Ru}_3\text{Cl}_4\text{S}(\text{C}_{18}\text{H}_{15}\text{P})_6] \cdot$
 $1.25\text{C}_4\text{H}_8\text{O} \cdot 0.7\text{H}_2\text{O}$
 $M_r = 2152.82$
 Triclinic, $P\bar{1}$
 $a = 14.183(4) \text{ \AA}$
 $b = 14.699(5) \text{ \AA}$
 $c = 25.814(9) \text{ \AA}$
 $\alpha = 91.904(11)^\circ$
 $\beta = 92.328(11)^\circ$
 $\gamma = 102.784(11)^\circ$
 $V = 5239(3) \text{ \AA}^3$

$Z = 2$
 $D_x = 1.365 \text{ Mg m}^{-3}$
 Synchrotron radiation,
 $\lambda = 0.5594 \text{ \AA}$
 Cell parameters from 1021
 reflections
 $\theta = 2.3\text{--}19.4^\circ$
 $\mu = 0.37 \text{ mm}^{-1}$
 $T = 123(2) \text{ K}$
 Prism, red
 $0.08 \times 0.03 \times 0.02 \text{ mm}$

Data collection

SMART 6000 detector mounted on
 a Bruker Kappa Platform
 φ scans
 Absorption correction: multi-scan
 (SADABS; Blessing, 1995;
 Sheldrick, 1996)
 $T_{\min} = 0.987$, $T_{\max} = 0.995$
 109758 measured reflections

15139 independent reflections
 12160 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.049$
 $\theta_{\max} = 19.0^\circ$
 $h = -14 \rightarrow 14$
 $k = -16 \rightarrow 16$
 $l = -29 \rightarrow 29$

Refinement

Refinement on F^2
 $R[F^2 > 2\sigma(F^2)] = 0.089$
 $wR(F^2) = 0.243$
 $S = 1.42$
 15139 reflections
 1156 parameters
 H atoms constrained OK?

$w = 1/[\sigma^2(F_o^2) + (0.06P)^2 + 6P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.001$
 $\Delta\rho_{\max} = 1.38 \text{ e \AA}^{-3}$
 $\Delta\rho_{\min} = -1.25 \text{ e \AA}^{-3}$

The asymmetric unit contains the complex cluster molecule (I) and four partially occupied tetrahydrofuran solvent sites. Additionally, there are four sites treated as partially occupied water O-atom sites, although they may be associated with further disordered and unresolved tetrahydrofuran sites. In general, the non-H-atom sites were modelled with anisotropic displacement parameters, and a riding atom model was used for the H atoms. The partially occupied non-H sites were modelled with isotropic displacement parameters. No H atoms were included in the model for the water sites.

Data collection: SMART-KAPPA (Bruker, 2001); cell refinement: SAINT (Bruker, 2001); data reduction: SAINT and XPREP (Bruker, 2001); program(s) used to solve structure: SHELXS97 (Sheldrick, 1997); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997); molecular graphics: TEXSAN for Windows (Molecular Structure Corporation, 1997), Xtal3.6 (Hall *et al.*, 1999), ORTEPII (Johnson, 1976) and WinGX (Farrugia, 1999); software used to prepare material for publication: SHELXL97.

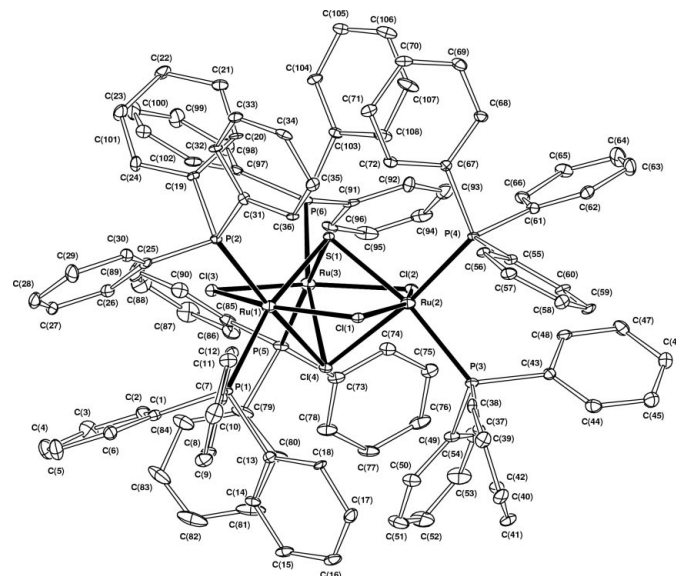


Figure 1
 ORTEPII (Johnson, 1976; Hall *et al.*, 1999) projection of (I), with displacement ellipsoids shown at the 20% probability level.

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References

- Blessing, R. H. (1995). *Acta Cryst.* **A51**, 33–38.
 Bruker (2001). SMART-KAPPA (Version 5.054), SAINT (Version 6.28A) and XPREP (Version 6.12). Bruker AXS Inc., Madison, Wisconsin, USA.
 Eckermann, A. L., Wunder, M., Fenske, D., Rauchfuss, T. B. & Wilson, S. R. (2002). *Inorg. Chem.* **41**, 2004–2006.
 Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
 Hall, S. R., du Boulay, D. J. & Olthoff-Hazekamp, R. (1999). Editors. *Xtal3.6 System*. University of Western Australia, Australia.
 Johnson, C. K. (1976). ORTEPII. Report ORNL-5138. Oak Ridge National Laboratory, Tennessee, USA.
 Lockemeyer, J. R., Rauchfuss, T. B. & Rheingold, A. L. (1989). *J. Am. Chem. Soc.* **111**, 5733–5738.
 Lockemeyer, J. R., Rauchfuss, T. B., Rheingold, A. L. & Wilson, S. R. (1989). *J. Am. Chem. Soc.* **111**, 8828–8834.
 Mashima, K., Hino, T. & Takaya, H. (1991). *Tetrahedron Lett.* **32**, 3101–3104.
 Molecular Structure Corporation (1997). TEXSAN for Windows. MSC, 3200 Research Forest Drive, The Woodlands, TX 77381, USA.
 Sheldrick, G. M. (1996). SADABS. University of Göttingen, Germany.
 Sheldrick, G. M. (1997). SHELXS97 and SHELXL97. University of Göttingen, Germany.